

Vocabulary: Solubility and Temperature



Vocabulary

- **Concentration** – a measure of how much of a given substance is mixed with another substance.
 - To measure the concentration of a *solution*, divide the mass of the *solute* by the volume of the *solvent*.
 - A common unit of concentration is grams per 100 milliliters (g/100 mL) or grams per deciliter (g/dL).
- **Dissolve** – to pass into solution.
 - For example, sugar or salt can dissolve into water. The resulting solution is a *homogeneous mixture*.
- **Homogeneous mixture** – a combination of two or more substances that is exactly the same throughout.
 - Any small sample of a homogeneous mixture would look exactly the same as any other sample, even at a microscopic level.
 - Solutions are homogeneous mixtures.
- **Solubility** – the maximum concentration of solute that can be dissolved in a solvent.
 - Solubility depends on the type of solute and solvent used and often depends on the temperature and pressure of the solvent.
- **Solubility curve** – a graph showing the relationship between solubility and another variable such as temperature or pressure.
- **Solute** – a substance that is dissolved in another substance to form a solution.
 - In salt water, the solute is salt.
- **Solution** – a homogeneous mixture of two or more substances.
 - Solutions generally consist of a solute that is dissolved into a solvent.
 - Solvents are generally liquids.
 - Solutes can be solids, liquids, or gases.
 - Examples of solutions include salt water, sugar water, and seltzer.
- **Solvent** – a solid, liquid, or gas in which a solute is dissolved to form a solution.
 - In salt water, the solvent is water.